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Work is found thethe frst law as $w = q - \Delta U$; thus $q = \Delta H$; $w = \Delta H - \Delta U$; 4. Isothermal Process Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for an isothermal process. These results also follow from the general results by using $dT = d(PV) = 0$ for an isothermal process.

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Chapter Thirteen. 13.1 a $Cu = 0.132$ a $Mg = 6.4 \times 10^{-4}$ 13.3 a $PbO = 0.134$ X $Cu = 0.018$. Increasing T decreases the extent to which Cu is removed. 13.5 a $C = 0.5pH^2 = 092$. atm 13.6 a $FeO = 9.9 \times 10^{-5}$ 13.7 (a) $ppHCO_2 = 2.15$, (b) a $C = 0.194$, (c) $P T = 5.16$ atm, (d) the total pressure.

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