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Making Esters

~~Grade 12 Esters~~

~~Practical Part 3~~

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~~Using Reflux and~~
~~Distillation~~
Preparation of
an Ester and
Separation by
Distillation -

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Reaction

Mechanism -
Carboxylic Acid
Derivatives

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setup-ester
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1: Synthesis of
an Ester from
Alcohol and

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3 Organic
Chemistry)

*Synthesis of an
ester*

*Esterification
Theory for
Practical Gr 12
Part 2:*

Synthesis of an
Ester from
Alcohol and
Carboxylic Acid

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Esters 4.

Organic

Preparation

\u0026

Purification of
an Ester. LDM2

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Comprehension

Exam Prep. *The
Fischer*

*Esterification:
Methyl*

Salicylate ~~HOW
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to make Practical
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(Odor and
Flavoring)

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Molecules:
Alcohols,
Carboxylic Acids
and Esters

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Mechanism:
making an ester
from a
carboxylic acid

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and an alcohol

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with Jay! Exp 9

Preparation of
Ethyl Ethanoate
Properties of
Ethanoic Acid:
Esterification
reaction (~~L-14~~)

~~Esterification~~
~~of R-COOH ||~~
~~with Mechanism~~

~~|| JEE NEET~~

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~~Synthesis Of~~

~~Ethyl benzoate~~

~~(sweet smelling~~

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The preparation
and the
practical report
will be
completed, in
lesson time,
under

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Supervision. The
preparation will
be conducted in
pairs, but
individual
reports will be
written. Stage 2
Chemistry
Practical (S):
PREPARATION OF
AN ESTER. The
preparation and
the practical

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Table 1

-Reagents for
preparation of
esters Add 4
drops of

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concentrated sulphuric acid to each test tube. Pour about 150 ml of tap water in a 250 ml beaker. Place the test tubes in the water and heat the water on a hot plate to a temperature of about 60°C - 75°C. Leave the

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Ester tubes in
the hot water
bath for 15
minutes.
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~~Preparation and
identification
of Esters—
Physical...~~

Part A:

Preparation of
Ester. Measure
15 mL of the
alcohol provided

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Preparation Of
and 10 mL
glacial (pure)
acetic acid into
the 50 mL pear-
shaped Quickfit
flask. Slowly
add about 1 mL
concentrated
sulfuric acid
and a few
boiling chips.
Assemble the
glassware for
reflux. Reflux

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the mixture for
30 minutes. Part
B: Isolation of
Pure Ester
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~~STAGE 2~~

~~CHEMISTRY~~

Practical

Investigation:

Preparation of
an Ester.

Response 1.

Skills checklist
(Teacher record)

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Part A: Teacher

observations:

Performing
reflux Clamped
correctly.

Vapours reached
the top of the
condenser. Use
of separating
funnel Careful.

Performing
distillation
Clamped
correctly.

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Adaptor joint
under strain.
Thermometer too
low.

~~Preparation of
Ester: Student
Report~~

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Synthesis,
Paper
Isolation, and
Purification of
an Ester

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Esters can be
synthesized
artificially in
labs by
combining
alcohols and
acids of
different
strengths. This
lab contained
the synthesis of
an ester by
allowing
different

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Carboxylic acids
to react with
alcohols, using
sulphuric acid
as a catalyst,
to create esters
and allow the
student to
identify them by
their distinct
smell.

~~Ester~~

~~Preparation Lab~~

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Assessment 5 (b)

The preparation of an organic liquid compound (e.g. the preparation of an ester). It is also expected that candidates will be aware of the procedures

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involved in the
purification and
re-distillation
at an
appropriate
temperature of
the product
obtained. The
Preparation of
Ethyl Ethanoate.
Reaction

~~A Level~~

~~Chemistry~~ — ~~Home~~

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To make a small ester like ethyl ethanoate, you can gently heat a mixture of ethanoic acid and ethanol in the presence of concentrated sulphuric acid, and distil off the ester as soon as it is formed. This

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prevents the
reverse reaction
happening. It
works well
because the
ester has the
lowest boiling
point of
anything
present.

~~preparation of
esters
chemguide~~

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5. EXEMPLARS OF
PRACTICAL WORK
AS FORMAL
ASSESSMENT TASKS

TERM 1 :

PRACTICAL WORK
KNOWLEDGE AREA :
MATTER AND
MATERIALS 5.1
PREPARATION OF
ESTERS AND SMELL
IDENTIFICATION
INTRODUCTION
Esters have a

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Paper

very fruity
smell. Naturally
occurring esters
are found in
fruits. Esters
can be
synthesised by
the reaction of
a carboxylic
acid and an
alcohol.

~~PHYSICAL~~
~~SCIENCES~~

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its content is
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Conditions. Tes
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registered in
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No 02017289)
with its
registered

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office at 26 Red
Lion Square
London WC1R 4HQ.
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Esters |
Teaching
Resources~~

assessment tasks

338 PREPARATION

FILE - Grade 12

Doc Scientia

Method: 1.

Number the test

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ester from 1 – 6
and place it in
the test tube
rack. 2. Add the
correct amount
of the
substances to
the test tubes
as indicated in
the table. Test
tube Carboxylic
acid Alcohol 1 1
cm³ ethanoic
acid 1 cm³

Get Free
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Ester Practical
methanol 2 1 cm³
ethanoic acid 1
cm³ ethanol
Task 2
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