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Solution: OH OH OH 0.50 mmol mL ??

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$n = c \times V = 0.350 \text{ mol/L} \times 25.0 \text{ mL} = 8.75 \text{ mmol HCl(aq)}$
 $n = c \times V = 0.500 \text{ mol/L} \times 5.0 \text{ mL} = 2.5 \text{ mmol NaOH(aq)}$
 $V = 120 \text{ mL}$
2. (a) Given: $V \text{ HCl(aq)} = 25.00 \text{ mL}$; $[\text{HCl(aq)}] = 0.350 \text{ mol/L}$; $V \text{ NaOH(aq)} = 5.0 \text{ mL}$; $[\text{NaOH(aq)}] = 0.500 \text{ mol/L}$; Required: $[\text{H}^+(\text{aq})]$, pH
Analysis: Before titration, amount of

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H⁺ is: H(aq) H(aq) H(aq) 0.350 mmol
mL ++ +=x = nc V

Section 8.7: Acid–Base Titration

Tutorial 1 Practice, page 547

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Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.040)(0.010)}{(0.14)^2} = 0.020$

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Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20$ mol/L; $[H_2(g)] = 0.15$ mol/L; $[I_2(g)] = 0.090$ mol/L; $K = 0.020$ Required: Q ; direction of reaction to reach equilibrium Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.15)(0.090)}{(0.20)^2}$

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Q=0.34

Section 7.5: Quantitative Changes in
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