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weight and stoichiometric
mole and mass air-fuel
ratios for the Oklahoma gas
mole composition given in
Table 3.4. Solution.

Equation (3.2), $\text{CH}_4 + 2\text{O}_2 +$
 $2(3.76)\text{N}_2\text{CO}_2 + 2\text{H}_2\text{O} +$
 $2(3.76)\text{N}_2$ (3.2) shows that

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there are $2 + 2(3.76) = 9.52$
moles of air required for
complete combustion of each
mole of methane.

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