

# Access Free Chemquest 53 Answers

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53. ChemQuest 17. Name:

\_\_\_\_\_ Date: \_\_\_\_\_ Hour:

\_\_\_\_\_ Information:

Separating Charges. Examine  
Figure 1 below where there  
are three pairs of metal  
spheres that have different  
amounts of. charge on them.  
The spheres in diagram C are  
closer than the others

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because they have the strongest. attraction.

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Thermodynamics

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This video is designed to help you complete the molarity chemquest found in Unit 11.

## *Molarity*

118 ChemQuest 39 Name: \_\_\_\_\_

Date: \_\_\_\_\_ Hour: \_\_\_\_\_

Information: Molarity

Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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An element exists as 4 different isotopes. 4.35% have a mass of 49.9461, 83.79% have a mass of 51.9405, 9.50% have a mass of 52.9407 amu, and 2.36% have a mass of 53.9389 amu.

a. What is the average atomic mass for this element? b. What is the identity of this element?

*Chemistry: ChemQuest Skill Practice (Semester 1 ...*

*traditionally and ...*

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Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$

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be one of the options to  
accompany you in imitation  
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question 1 and four times  
your answer to question 2.  $4$   
 $+ 4(7) = 4 + 28 = 32$

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be one of the options to  
accompany you in imitation  
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15. Given your answer to  
question 14 and the reaction  
you wrote in question 13,  
calculate the molarity of  
the nitrate ions ( $\text{NO}_3^-$ ) in  
the solution. Title:

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(Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.)  
12 moles. For every mole of propane that combusts 3 moles of CO<sub>2</sub> are produced, so the number of moles of CO<sub>2</sub> that can be produced when 4 moles of propane combusts =  $4(3) = 12$ .

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