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22. The concentration of either
the H^+ ion or the OH^- ion is
given for four aqueous solutions
at 298 K. For each solution,
calculate $[H^+]$ or $[OH^-]$. State
whether the solution is acidic,

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MULTIPLE CHOICE QUESTIONS
NOTE: A table of standard
reduction potentials is required to
work many of these problems. 1.
Complete and balance the
following redox equation. When

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properly balanced using the
smallest whole-number
coefficients, the coefficient of S is
 $\text{H}_2\text{S} + \text{HNO}_3 \rightarrow \text{S} + \text{NO}$ (acidic
solution) 2.

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