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PART I In the space provided, write the
letter of the term or phrase that best
completes each statement or best answers
each question. 1. Knowing the mole ratio
of a reactant and product in a chemical

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Answers would allow you to determine a. the energy released in the reaction. b. the speed of the reaction.

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Chapter Review Answers 0.0209 mol

AgNO₃ Aspirin production a. 13.5 kg

C₉H₈O₄ b. 7.66 kg C₄H₆O₃ c. 4.29 L

HC₂H₃O₂ Ideal stoichiometry

calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Input it if you want to ...

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Answer Key Answer: 4.93×10^{-5} L or $49.3 \mu\text{L}$ In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we

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Answers can determine which reactant is limiting by comparing the. Page 5/28.

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2 of 2) 19. The mass of element X found in
1.00 mole of each of four different
compounds is 28.0 g, 42.0 g, 56.0 g, and
70 g, respectively. The possible atomic

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weight of X is a. 8.00 b. 14.0

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