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As liquids absorb heat at their boiling points, the temperature remains constant

while they vaporize. true _____ Chapter 17

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(continued) Use the heating curve for water
shown below to answer Questions 5, 6, and 7.

section 17.1 the flow of energy heat and work

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0.100 g of H₂ and an excess of O₂ are
compressed into a bomb calorimeter containing
1200 g of water. The temperature before the
reaction is 25.00 °C, and after the reaction
it goes to 27.16 °C.

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CHAPTER 17, Thermochemistry (continued) Use the heating curve for water shown below to answer Questions 5, 6, and 7. Heating Curve for Water Boiling point ~: :>,-----, 1 ~ Melting E point ~ <--_-----'I Heat supplied 5. Label the melting point and boiling point temperatures on the graph. 6. What happens to the temperature during melting and vaporization?

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Chapter 17 Thermochemistry 187 10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520-526) This section explains heat

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transfers that occur during melting, freezing, boiling, and condensing.

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)

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1. How much energy must be absorbed by 20.0 g of water to increase its temperature from 283.0 °C to 303.0 °C? 2. When 15.0 g of steam drops in temperature from 275.0 °C to 250.0 °C, how much heat energy is released?

Thermochemistry Problems - Worksheet Number One

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