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Rate Laws –

Chemistry Review –

Order of Reaction

\u0026 Equations

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Part 1 Chapter 14 –

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questions of chemical

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Reactions

Kinetics: Initial Rates
and Integrated Rate

Laws **The Rate Law**

*Chapter 16 – Acid-
Base Equilibria: Part 1
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Chemistry 11.4

*Factors that affect
Reaction Rate*

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probable Question

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Instantaneous Rate •

We can plot [C. 4. H.

9. Cl] versus time. •

The rate at any
instant in time is
called the .

instantaneous rate. •

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It is the slope of the straight line tangent to the curve at that instant. •

Instantaneous rate is different from average rate.

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Chemical Kinetics

14.4. Use the equations in the AP Chemistry test booklet to work kinetics

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problems. Explain the concept of reaction half-life and describe the relationship between half-life and rate constant for a first-order reaction. Use graphical analysis to determine whether the rate law for a reaction is first or second order.

14.5-14.7

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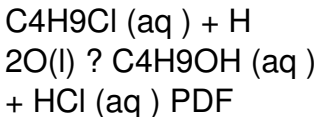
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Reaction Rates • All reactions slow down over time. •

Therefore, the best indicator of the rate of a reaction is the instantaneous rate near the beginning.



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14.1: Factors that Affect Reaction Rates. chemical kinetics – area of chemistry dealing with speeds/rates of reactions. rates of

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reactions affected by
four factors.
concentrations of
reactants.

temperature at which
reaction occurs.

presence of a
catalyst. surface area
of solid or liquid
reactants and/or
catalysts.

14.S: Chemical Kinetics (Summary)

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-Chemistry

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Question 1: A catalyst lowers the activation energy of a reaction from 20 kJ mol^{-1} to 10 kJ mol^{-1} . The temperature at which the catalyzed reaction will have the same rate as that of the uncatalyzed reaction at 27°C is 327°C . $32.7^\circ \text{C} + 23^\circ \text{C}$.

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Correct Option is :

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CHOICE. Choose the

one alternative that

best completes the

statement or answers

the question. ... 14)

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The rate law for a reaction is $\text{rate} = k[A][B]^2$... The graph shown below depicts the relationship between concentration and time for the following chemical reaction. The slope of this ...

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Chemistry Chapter 14. Chemical Kinetics

- 7 - • Consider the reaction: $\text{NH}_4^+(\text{aq}) + \text{NO}_2^-(\text{aq}) \rightarrow \text{N}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ • We measure initial reaction rates. • The initial rate is the instantaneous rate at time $t = 0$. • We find

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this at various initial concentrations of each reactant. • As $[\text{NH}_4^+]$ doubles with $[\text{NO}_2^-]$ Chapter 14. Chemical Kinetics AP Chemistry Chapter 14.

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The slope of this ...

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Major topics:

integrated

zero/first/second

order rate laws,

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zero/first/second

order reactions

graphically, & half-life

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It has been found that for a chemical reaction with rise in temperature by 10°C , the rate constant gets nearly doubled.

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15. The temperature coefficient of a reaction is the ratio of the rate constants of the reaction at two temperatures differing from one another by 10°C .

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